Straightforward One-Pot Syntheses of Silylamides of Magnesium and Calcium via an In Situ Grignard Metalation Method

Sven Krieck
Philipp Schüler
Jan M. Peschel
Matthias Westerhausen*

Friedrich Schiller University Jena, Institute of Inorganic and Analytical Chemistry, Humboldtstraße 8, D-07743 Jena, Germany
m.wel@uni-jena.de

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Abstract Calcium bis[bis(trimethylsilyl)amide] \((\text{Ca(HMDS)}_2)\) is a widely used reagent in diverse stoichiometric and catalytic applications. These processes necessitate a straightforward and large-scale access of this complex. Calcium does not react with primary and secondary amines, but the addition of excess bromoethane to a mixture of calcium turnings and amines in THF at room temperature yields the corresponding calcium bis(amides), calcium bromide and ethane. This in situ Grignard metalation method (iGMM) allows the preparation of calcium bis(amides) from secondary and primary trialkylsilyl-substituted amines and anilines on a multigram scale.

1 Background

The development of diverse procedures for the synthesis of calcium bis[bis(trimethylsilyl)amide] \((\text{Ca(HMDS)}_2)\) is based on the desperate need of soluble organocalcium complexes for various applications. Therefore, different strategies were developed independently at the beginning of the 1990s. The first syntheses of calcium bis[bis(trimethylsilyl)amide] \((\text{Ca(HMDS)}_2)\) were performed via transmetalation of Hg(HMDS)\(_2\) or Sn(HMDS)\(_2\) via metathetical approaches of AHMDS \((A = \text{alkali metal})\) with calcium alkoxides\(^3\) or pseudohalides such as trifluoromethanesulfonates (triflates)\(^4\) and arenesulfonates.\(^5\) The metathesis reaction of KHMDS with CaI\(_2\) in tetrahydrofuran\(^6,7\) and toluene\(^8\) yields thf adducts and unsolvated Ca(HMDS)\(_2\), respectively. Direct metalation of \((\text{H})\text{HMDS}\) with calcium is very challenging and requires activation of the alkaline earth metal via metal vapor synthesis and co-condensation with the organic solvent.\(^3\) Alternatively, ammonia-saturated solvents can be used for the direct metalation of \((\text{H})\text{HMDS}\) with calcium turnings.\(^9\) These procedures are depicted in Scheme 1.

All these procedures are related to severe drawbacks. Transmetalation protocols require the preparation, isolation and distillation of the precursor organometallics (Hg(HMDS)\(_2\) or Sn(HMDS)\(_2\)). Metathetical approaches necessitate an exact stoichiometry to avoid formation of calcicates of the type \([\text{Ca(HMDS)}_3]^-\) (excess of AHMDS) and halide-containing product mixtures (substoichiometric AHMDS). The Ca(HMDS)\(_2\) product commonly contains at least traces of potassium (potassium calciate formation...
and/or traces of potassium iodide) that have to be removed by repeated recrystallization efforts. Prior to the direct metalation of (H)HMDS the calcium metal has to be activated to apply pyrophoric metal powders. In our hands, calcium pieces are sluggish in reactions in ammonia-saturated solutions and dull, grayish reaction mixtures are obtained which prohibit the isolation of a pure product. These initial investigations are summarized and evaluated in more detail elsewhere.10

Due to these reasons, diverse strategies have been developed to accelerate the synthesis of highly pure Ca(HMDS)2. The reaction of calcium with (H)HMDS in THF can be ensured via addition of BiPh3 under ultrasonication.11 The organometallic metalation of (H)HMDS with dibenzylcalcium12 circumvents the necessity of metal activation prior to use. Arylcalcium reagents (heavy Grignard reagents)13 are easily available and represent suitable metalation reagents for the syntheses of calcium amides.14

Drawbacks of these procedures are obvious because another metalorganic precursor is required to prepare the metalation reagent.

The bis(trimethylsilyl)amido ligands guarantee not only solubility of the respective metal complexes in common organic solvents, but variations of this anion allow modification of the electronic and (especially) steric properties. Transmetalation of Sn[N(SiMe2CH2)2]2 with calcium in THF yields the corresponding tris(thf) adduct of calcium bis(2,2,5,5-tetramethyl-1-aza-2,5-disilylcyclopentanide).15 In contrast, enlargement of the amido ligand and the use of Sn[N(SiMe3)Si(SiMe3)3]2 in the transmetalation procedure only gives the aminyl radicals and finally the amine HN(SiMe3)Si(SiMe3)3, but a calcium bis(amide) cannot be isolated. Suitable procedures for the preparation of Ca[N(SiHMe2)2]2 are the metathetical approach with KN(SiHMe2)2 and CaI2 as well as the transamination of Ca(HMDS)2 with HN(SiHMe2)2.16 The metathesis reaction of the potassium amides with calcium iodide offers access to solvated bulky calcium bis(amides) with amido ligands like

Biographical Sketches

Sven Krieck (left) studied chemistry at Friedrich Schiller University Jena, Germany, and graduated in 2007 with his diploma thesis in the field of electrochemical studies on heterosupramolecular aggregates in the group of Professor Günter Kreisel. He then joined the group of Professor Matthias Westerhausen and completed his Ph.D. in April 2010 working on the stabilization of organocalcium compounds. His thesis was awarded with the university prize as best thesis in 2010. After a postdoctoral fellowship with Professor Nadia Mösch-Zanetti at Karl Franzens University in Graz, Austria, he returned to the Westerhausen group in 2011. Today he is lecturer at the Institute of Inorganic and Analytical Chemistry at Friedrich Schiller University Jena, Germany.

Philipp Schüler (second from left) studied the synthesis and coordination chemistry of alkaline earth metal bis(bis(trimethylsilyl)amides) at Friedrich Schiller University Jena, Germany, and obtained his M.Sc. degree in 2018. He is currently a Ph.D. student in the Westerhausen group at the same institution.

Jan M. Peschel (second from right) studied chemistry at Friedrich Schiller University Jena, Germany, and obtained his B.Sc. degree in 2017. During his laboratory courses he investigated the synthesis and coordination chemistry of alkaline earth metal bis(bis(trimethylsilyl)amides).

Matthias Westerhausen (right) obtained his diploma degree in chemistry in 1983 from Philipps University in Marburg, Germany, and studied for his Ph.D. thesis at the University of Stuttgart, Germany, under the supervision of Professor Gerd Becker. In 1987/88, he worked as a postdoctoral fellow with Professor Robert T. Paine at the University of New Mexico in Albuquerque, USA, in the field of phosphanylboranes. Back at the University of Stuttgart, he finished his habilitation in the Institute of Inorganic Chemistry in December 1994 and received the venia legendi for Inorganic Chemistry in February 1995. From 1996 to 2004 he was a professor at Ludwig Maximilians University Munich where he was also vice rector from 2001 to 2004. He was awarded the Teaching Excellence Award of the State of Bavaria in 1998. Since 2004, he has been teaching and researching at Friedrich Schiller University Jena, Germany.
N(SiMe3)2SiMe2tBu, N(SiMe3)2SiPh2tBu), and N(SiMe3)2SiPh3.17 The salt metathesis reaction even allows the synthesis of unsolvated calcium bis[(trimethylsilyl)amide] with a nearly linear N–Ca–N fragment of 172.62(11)°.18 This metathetical approach also provides access to substituted N-trimethylsilylanilides with bulky aryl groups like 2,4,6-trimethylphenyl (mesityl, Mes)19 and 2,6-diisopropylphenyl (Dipp).20

Furthermore, the co-ligands of Ca(HMDS)2 can be varied to also influence the shielding of the reactive Ca–N bonds. Structural studies have been reported for dimeric [Ca(HMDS)2]n21 as well as for the ether solvates [(dme)Ca(HMDS)2]21 and [(thf)2Ca(HMDS)2].22 Electroneutral co-ligands with nitrogen donor atoms like tmeda24 and imidazole25 form stable complexes too. N-Heterocyclic carbenes are strong σ-donors and as such they are strong ligands for complexation of Ca(HMDS)2, comparable to hexamethylphosphoric acid triamide.26

Since the discovery of these Ca(HMDS)2 complexes in the early 1990s, multiple application fields have been established such as materials science, as ligand transfer and metalation reagents, as precursors in hydrofunctionalization reactions, and as polymerization initiators. This demand justifies an ongoing interest in these calcium complexes leading to a vast development of improved preparative protocols and expansion of the fields of application. Beneficial properties of this alkaline earth metal are based on its position in the periodic table between the alkaline metals (highly electropositive metal, ionic Ca–N bonds) and the early transition metals (d-orbital participation, highly Lewis acidic ions, catalytic activity). Other useful properties include non-toxicity, global abundance, and the inexpensive nature of this alkaline earth metal.

Due to these reasons, we have developed a straightforward strategy for the synthesis of Ca(HMDS)2 that allows the production of a large amount of this reagent without the necessity to prepare a metalorganic precursor. Furthermore, this method is not limited to the bis(trimethylsilyl)amides, but also allows the synthesis of other as yet unknown calcium bis(amides).

2 The In Situ Grignard Metalation Method (iGMM)

Calcium readily forms heavy Grignard reagents via the direct synthesis of activated calcium metal (Ca*) with iodo- and bromoarenes or iodomethyl- and bromomethyl-trimethylsilane yielding ArCaX or Me5SiCH2CaX (X = Br, I). However, aliphatic hydrocarbylcalcium halides (RCaX) are not accessible via this approach (with the exception of cyclopropylcalcium halide).27 Instead, Wurtz-type C–C coupling reactions dominate and alkanes R–R form nearly quantitatively. The reason for this strikingly different behavior of aromatic and aliphatic hydrocarbyl iodides and bromides R–X (X = Br, I) lies in the nature of the hydrocarbyl halide radical anions after the first electron transfer from the metal onto R–X. Without stabilization of the negative charge (via delocalization into the aryl group or via negative hyperconjugation to the trimethylsilyl groups) the C–X bond is elongated much more, easing dissociation into the halide anion and the alkyl radical.

Calcium is unable to directly metalate (H)HMDS in common organic solvents (aliphatic and aromatic hydrocarbons, ethers) regardless of a prior activation procedure. Therefore, commercially available calcium granules and this amine were placed together in a Schlenk flask in THF. Next, ethyl bromide, dissolved in THF, was added dropwise at room temperature and the reaction mixture was stirred for an additional three hours to complete the reaction. This reaction is summarized in Scheme 2.

The straightforward synthesis of Ca(HMDS)2 in THF using the in situ Grignard metalation method (iGMM)

The calcium turnings slowly dissolve and evolution of ethane was observed. Equimolar amounts of calcium bromide and Ca(HMDS)2 formed with the Schlenk-type equilibrium in favor of the homoleptic compounds. This finding is in agreement with earlier studies.24 This procedure also worked nicely with magnesium turnings, however, in this case, the equilibrium lies completely on the side of the heteroleptic complex (HMDS)MgBr, an important Hauser base with strong metalating power (Scheme 3). Addition of 1,4-dioxane leads to an equilibrium between the Hauser base on the one side and homometallic Mg(HMDS)2 and magnesium bromide on the other (see the Supporting Information). In the past, these Hauser bases had to be prepared via metalation of bis(trimethylsilyl)amine with previously prepared alkylmagnesium bromide. Alternatively, homoleptic Mg(HMDS)2 can be prepared by an organometallic metalation of (H)HMDS with commercially available dibutylmagnesium, thereby avoiding formation of the Hauser base reagent.

The results with respect to the reaction conditions are summarized in Table 1. An equimolar amount of calcium granules and bromoethane leads to a conversion of only 43% (entry 1) because ether cleavage (formation of R–H)
and Wurtz-type coupling reactions (formation of R–R) compete with the deprotonation of (H)HMDS. If another half equivalent of EtBr is added after two hours almost complete conversion is obtained (entry 2). Activation of the calcium metal and the use of pyrophoric activated calcium powder accelerates the reaction and higher yields are achieved (entries 3 and 4). The use of iodoethane gave a very similar yield (entry 5). Larger alkyl groups lower the yields significantly, whereas the calculation of (H)HMDS cannot be initiated by a chloroalkane (entry 7). Benzyl bromide, too, is a suitable initiator (entry 12). Contrary to this finding, bromobenzene shows no activation properties in agreement with earlier studies, with aryl halides only reacting with acetylbenzene shows no activation properties in agreement with earlier studies, with aryl halides only reacting with acetylbenzene shows no activation properties in agreement with earlier studies, with aryl halides only reacting with acetylbenzene shows no activation properties in agreement with earlier studies.

The results are listed in Table 2 together with the standardized reaction conditions. In all cases, 1.5 equivalents of calcium turnings were used and combined with 1 equivalent of the amine in THF. Then bromoethane was added dropwise over a period of 30 minutes at room temperature. The presence of one trialkylsilyl substituent on the nitrogen atom ensures a smooth reaction with high yields, regardless of the bulkiness of the trialkylsilyl group (entries 1 and 4–6). Also, mixed alkyl-trimethylsilylamines (entry 5) represent a suitable substrate under these reaction conditions. Anilines can also be calcinated under these conditions (entry 3), however, the yield is lower and longer reaction times are required. Dialkylamines (entry 2) immediately form ammonium bromides during the addition of bromoethane, hampering the calciation reaction. For the same reason, primary and secondary phosphines are not calcinated under these reaction conditions, but phosphonium bromides are formed during the addition of bromoethane.

In summary, aryl- (anilides) and trialkylsilyl-substituted amines of calcium are accessible via this novel iGMM. Commercially available calcium granules are a suitable substrate for the calciation of these primary and secondary...
amines. In contrast to the calcium-based procedure, Hauser bases form if magnesium turnings are used in this reaction. Due to competing Wurtz-type coupling reactions a slight excess of calcium is beneficial. This synthesis of Ca(HMDS)$_2$ can easily be upscaled to multigram scale without preceding elaborate and time-consuming preparations (such as purification and activation of calcium, preparation of calcination reagents or dealing with ammonia-saturated ethereal solvents).

The results are depicted in Scheme 5. Direct metalation of (HMDS) and other amines with calcium in THF is impossible. Addition of bromoethane to this mixture at room temperature initiates a gentle and smooth deprotonation and formation of Ca(HMDS)$_2$. Ethane escapes from this reaction and calcium bromide precipitates without disturbance of the ongoing calcination process. Heteroleptic (HMDS)CaBr cannot be observed whereas the use of magnesium turnings favors the formation of the Hauser base (HMDS)MgBr. Addition of 1,4-dioxane leads to an equilibrium between the heteroleptic Hauser base on the one side and the homoleptic complexes Mg(HMDS)$_2$ and MgBr$_2$ on the other.

The thus prepared Ca(HMDS)$_2$ is free of other metals, but the solution contains calcium bromide which does not interfere with the reactivity of the amide because Hauser-base-like products have not been observed in these ethereal solutions. Nevertheless, removal of calcium bromide can easily be achieved. For this purpose, all volatile materials are removed in vacuo and the remaining residue is extracted with pentane. Calcium bromide is insoluble in alkanes whereas the calcium amides precipitate as colorless crystals upon cooling of this alkane solution.

In summary, the iGMM is an extremely valuable and straightforward procedure to prepare quantitatively, on a multigram scale, new calcium bis(amides) with aryl or tris(hydrocarbyl)silyl substituents at the nitrogen atom and, also for the preparation of heteroleptic amidomagnesium bromide (Hauser base) solutions. The magnesium-based Hauser bases and their calcium-based congeners can be used as strong metalation reagents and furthermore, amidomagnesium halides containing stoichiometric amounts of LiCl are known as Turbo-Hauser bases.$^{28}$

### Properties of [(thf)$_2$M(HMDS)$_2$]

The bis(trimethylsilyl)amides of the alkaline earth metals magnesium, calcium, strontium, and barium form bis(thf) adducts with distorted tetrahedral coordination spheres of the metal centers.$^{1,2,29}$ These bis(thf) adducts [(thf)$_2$M(HMDS)$_2$] show characteristic NMR parameters that are summarized in Table 3.$^{2,10b}$ With increasing electronegativity of the central metal atom, the resonances of the $^{29}$Si nucleus show a high-field shift due to back donation of negative charge from the nitrogen atom into the $\sigma^*$ bond, a property known as negative hyperconjugation. If only one trimethylsilyl group is bound at the nitrogen atom, then this effect is enhanced leading to increased hyperconjugative effects and an even stronger high-field shift of the $^{29}$Si resonance as observed for [(thf)$_2$Ca(N(pBu)(SiMe$_3$)$_2$)]$_2$ with a value of $\delta = -20.7$. The $\delta^{(15N)}$ values of the [15N]-labeled complexes also strongly depend on the size and hardness of the alkaline earth metal.

| $^{29}$Si resonance as observed for [(thf)$_2$Ca(N(pBu)(SiMe$_3$)$_2$)]$_2$ with a value of $\delta = -20.7$. The $\delta^{(15N)}$ values of the [15N]-labeled complexes also strongly depend on the size and hardness of the alkaline earth metal. |
|---|---|---|---|---|
| $\delta^{(15N)}$ | Mg | Ca | Sr | Ba |
| $\delta^{(15N)}$ | 35.45 | -304.5 | -301.9 | -275.9 |
| $\delta^{(15N)}$ | -8.6 | -15.0 | -16.7 | -18.4 |
| $\delta^{(15N)}$ | 8.1 | 9.0 | 9.1 | 9.9 |

The molecular structures of the [(thf)$_2$M(HMDS)$_2$] complexes are very much alike with characteristic trends dependent on the size and electronegativity of the alkaline earth metal. Selected values are listed in Table 4. As expected, the M–N and M–O bond lengths increase with the radius of the metal. Decreasing intramolecular repulsion with increasing size of the metal center is evident from the decreasing N–M–N bond angles. The molecular structures of the [(thf)$_2$M(HMDS)$_2$] complexes are very much alike with characteristic trends dependent on the size and electronegativity of the alkaline earth metal. Selected values are listed in Table 4. As expected, the M–N and M–O bond lengths increase with the radius of the metal. Decreasing intramolecular repulsion with increasing size of the metal center is evident from the decreasing N–M–N bond angles. This interaction leads to large Si–N–Si bond angles.

Due to steric shielding of the reactive M–N bonds by the bulky HMDS ligands, these complexes show only a moderate sensitivity toward air and hence, isolated crystalline
materials have also been prepared and functionalized using simple and straightforward synthesis of [(thf)2M(HMDS)2]. The nature of Ca(HMDS)2 accelerates its use for the synthesis of [(thf)2M(HMDS)2] complexes of magnesium 1 and calcium, 2 acetylenation of carbodiimides.42 Inorganic–organic hybrid processes, Ca(HMDS)2 is a valuable and easy-to-handle prebiopolymers such as polylactides. 40 For diverse catalytic metallocenes,33 and alkynyl complexes. 34 Lewis acid–base ples are discussed below.

The complexes [(thf)2M(HMDS)2] (M = Mg, Ca) are valu-able synthons in metalation reactions of acidic compounds; early reactivity studies include examples of the synthesis of phosphinides,30 thiolates,31 selenolates,32 tellurolates,32 metallocenes,33 and alkynyl complexes.34 Lewis acid–base reactions with trialkylalanes 35 and -gallanes 36 allowed the synthesis of Ca–C bonds to bridging alkyl groups. Very recently, dimethylcalcium was prepared via a ligand exchange between Ca(HMDS)2 and methylthiium.37 Addition of [(thf)2Ca(HMDS)2] to nitriles allows the synthesis of amidi-nates with high yields.38 Recently, these compounds also allowed the synthesis of calcium hydride cages.39

Besides these stoichiometric conversions many catalytic applications have recently been elucidated. The non-toxic nature of Ca(HMDS)2 accelerates its use for the synthesis of biopolymers such as polyacetylides.40 For diverse catalytic processes, Ca(HMDS)2 is a valuable and easy-to-handle precatalyst. Representative examples include the calcium-mediated hydroamination of unsaturated compounds (like alkenes, alkynes, heterocumulenes)41 as well as the hydroacetylenation of carbodiimides.42 Inorganic–organic hybrid materials have also been prepared and functionalized using this compound as a catalyst.43

These selected examples demonstrate the necessity of a simple and straightforward synthesis of [(thf)2M(HMDS)2] (M = Mg, Ca) with high yields that can be safely performed without the requirement of sophisticated preparative skills. This novel GMM fulfills all preconditions for an easy access to widely use these environmentally benign reagents.

All manipulations were carried out under an inert nitrogen atmosphere using standard Schlenk techniques unless otherwise noted. All substrates were purchased from Alfa Aesar, abcr, Sigma Aldrich, or TCI and were used without further purification. The solvents were dried over KOH and subsequently distilled over sodium/benzophenone under a nitrogen atmosphere prior to use. Deuterated solvents were dried over sodium, distilled, degassed, and stored under nitrogen over sodium. HN(tBu)(SiMe3)40 and iPr3SiNH43 were prepared according to literature protocols. The yields given are crystalline yields and not optimized. IR spectra were recorded on a Bruker Alpha spectrometer. 1H, 29Si and 13C(1H) NMR spectra were recorded on a Bruker Avance III 400 spectrometer. Chemical shifts are reported in parts per million relative to SiMe4 as an external standard referenced to the residual proton signal of the solvent. The purity of the compounds was verified by NMR spectroscopy.

<table>
<thead>
<tr>
<th>[(thf)2MgBr(HMDS)]</th>
<th>[(thf)2Ca(HMDS)2]</th>
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<tr>
<td>Magnesium (0.97 g, 40.4 mmol, 1.1 equiv) was suspended in THF (45 mL) and (H)HMDS (5.9 g, 36.4 mmol) was added in one portion. EtBr (2.7 mL, 36.4 mmol, 1 equiv) was added dropwise over 30 min yielding a clear slightly yellow solution which can be used as a stock solution for this Hauser base (0.87 M, quant.). Excess magnesium metal was removed by filtration. 1H NMR (400 MHz, C6D6, 296 K): δ = 3.73 (m, 8 H, thf), 1.44 (m, 8 H, thf), 0.23 (s, 18 H, SiMe3). 13C(1H) NMR (101 MHz, C 6D6, 296 K): δ = 69.7 (thf), 24.8 (thf), 5.9 (SiMe3).</td>
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<tr>
<td>[(thf)2Ca(HMDS)2]</td>
<td>Calcium (2.0 g, 50 mmol, 1.1 equiv), (H)HMDS (9.6 mL, 45.4 mmol, 1 equiv) and THF (45 mL) were placed in a Schlenk flask. EtBr (3.4 mL, 45.4 mmol, 1 equiv) in THF (5 mL) was added dropwise while a moderate gas evolution was observed. The slurry was stirred for 2 h, a second amount of EtBr (2 mL, 26.8 mmol, 0.6 equiv) was added in one portion and the mixture was stirred for 2 h at room temperature. The solution was decanted and could be used after acidic titration (0.62 M) as a stock solution in organic synthesis (Br- concentration: 0.17 mol/l).</td>
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<td>The crystalline compound [(thf)2Ca(HMDS)2] was isolated as follows: The solvents were removed under reduced pressure and the residue redissolved in a mixture of n-pentane/THF (1 × 40 mL:1 mL) and filtered through a frit covered with diatomaceous earth (or separated via decantation/centrifugation). The filtrate was stored at −40 °C yielding solid [(thf)2Ca(HMDS)2] as colorless and halide-free crystals which were collected and dried in vacuo (6.2 g, 12 mmol, 49%).</td>
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<td>The analytical data are in accordance with the literature parameters.1,2</td>
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**Table 4** Selected Structural Parameters (Average Values, Bond Lengths [pm] and Angles [°]) of the Alkaline Earth Metal [(thf)2M(HMDS)2] Complexes

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<td>259.2</td>
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4 Applications and Perspective

In the initial 15 years after the first preparation of these [(thf)2M(HMDS)2] complexes of magnesium1 and calcium,2 the physical and chemical properties have been elucidated in detail. For approximately 20 years, widespread stoicho-metric and catalytic applications have been identified and explored, especially for these environmentally benign magnes-nium and calcium complexes. A few representative exam-ples are discussed below.

The complexes [(thf)2M(HMDS)2] (M = Mg, Ca) are valu-able synthons in metalation reactions of acidic compounds; early reactivity studies include examples of the synthesis of phosphinides,30 thiolates,31 selenolates,32 tellurolates,32 metallocenes,33 and alkynyl complexes.34 Lewis acid–base reactions with trialkylalanes35 and -gallanes36 allowed the synthesis of Ca–C bonds to bridging alkyl groups. Very recently, dimethylcalcium was prepared via a ligand exchange between Ca(HMDS)2 and methylthiium.37 Addition of [(thf)2Ca(HMDS)2] to nitriles allows the synthesis of amidi-nates with high yields.38 Recently, these compounds also allowed the synthesis of calcium hydride cages.39

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These selected examples demonstrate the necessity of a simple and straightforward synthesis of [(thf)2M(HMDS)2] (M = Mg, Ca) with high yields that can be safely performed without the requirement of sophisticated preparative skills. This novel GMM fulfills all preconditions for an easy access to widely use these environmentally benign reagents.
of EtBr (1.4 mL, 18 mmol, 1 equiv) was added in one portion and the reaction mixture was stirred overnight at room temperature. After decanting of the supernatant, one aliquot of the solution was titrated by acidimetric titration (0.51 M). The solvent was removed under reduced pressure, the residue was redissolved in with n-hexane (40 mL) and filtered through a frit covered with diatomaceous earth. The clear filtrate was stored at −20 °C and the obtained precipitate was collected and dried in vacuo yielding the amide as a colorless powder (2.6 g, 5.5 mmol, 31%).

**IR (ATR):** 2938, 2887, 1460, 1347, 1235, 1194, 1056, 1029, 991, 917, 853, 816, 752, 687, 651, 611, 504, 476, 432 cm⁻¹.

**13C NMR (101 MHz, C₆D₆, 296 K):**

δ = 176.5, 138.0, 132.1, 129.8, 129.2, 128.8, 128.2, 125.1, 123.7, 123.2, 120.7, 117.6, 114.7, 113.5, 110.4, 55.7, 46.8, 45.6, 44.8, 44.6, 43.8, 43.6, 41.2, 38.8, 38.4, 38.0, 37.8, 37.6, 37.4, 37.2, 36.8, 36.6, 36.4, 36.2, 35.9, 35.7, 35.5, 35.3, 35.1, 34.8, 34.6, 34.4, 34.2, 34.0, 33.8, 33.6, 33.4, 33.2, 33.0, 32.8, 32.6, 32.4, 32.2, 32.0, 31.8, 31.6, 31.4, 31.2, 31.0, 30.8, 30.6, 30.4, 30.2, 30.0, 29.8, 29.6, 29.4, 29.2, 29.0, 28.8, 28.6, 28.4, 28.2, 28.0, 27.8, 27.6, 27.4, 27.2, 27.0, 26.8, 26.6, 26.4, 26.2, 26.0, 25.8, 25.6, 25.4, 25.2, 25.0, 24.8, 24.6, 24.4, 24.2, 24.0, 23.8, 23.6, 23.4, 23.2, 23.0, 22.8, 22.6, 22.4, 22.2, 22.0, 21.8, 21.6, 21.4, 21.2, 21.0, 20.8, 20.6, 20.4, 20.2, 20.0, 19.8, 19.6, 19.4, 19.2, 19.0, 18.8, 18.6, 18.4, 18.2, 18.0, 17.8, 17.6, 17.4, 17.2, 17.0, 16.8, 16.6, 16.4, 16.2, 16.0, 15.8, 15.6, 15.4, 15.2, 15.0, 14.8, 14.6, 14.4, 14.2, 14.0, 13.8, 13.6, 13.4, 13.2, 13.0, 12.8, 12.6, 12.4, 12.2, 12.0, 11.8, 11.6, 11.4, 11.2, 11.0, 10.8, 10.6, 10.4, 10.2, 10.0, 9.8, 9.6, 9.4, 9.2, 9.0, 8.8, 8.6, 8.4, 8.2, 8.0, 7.8, 7.6, 7.4, 7.2, 7.0, 6.8, 6.6, 6.4, 6.2, 6.0, 5.8, 5.6, 5.4, 5.2, 5.0, 4.8, 4.6, 4.4, 4.2, 4.0, 3.8, 3.6, 3.4, 3.2, 3.0, 2.8, 2.6, 2.4, 2.2, 2.0, 1.8, 1.6, 1.4, 1.2, 1.0, 0.8, 0.6, 0.4, 0.2, 0.0.

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**Supporting Information**

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**References**


