Oxidative β-Halogenation of Alcohols: A Concise and Diastereoselective Approach to Halohydrins

Lingsheng Aia, Weijin Wanga, Jialiang Weia, Qing Lia, Song Song*a,b, Ning Jiao*a,b

Abstract β-Halohydrins bearing transformable halo- and hydroxyl groups, are easily converted into various valuable blocks in organic and pharmaceutical synthesis. A diastereoselective β-halogenation of benzylic alcohols was achieved under simple and low-cost conditions, which provided a direct synthesis of β-halohydrins. The simple reaction conditions, easily available reagents, high diastereoselectivities, and additional oxidant-free make this reaction very attractive and practical.

Key words halogenation, alcohols, dimethyl sulfoxide, halohydrins, oxidation

Organohalides are one of the most widespread and important chemicals and are present in more than 4500 natural products, as well as a great number of industrially valuable products such as pharmaceuticals, fire retardants, agrochemicals, and some new materials. In addition, it is no doubt that organohalides with their general reactivity make the chemical synthesis more simple, accessible, and valuable. β-Halohydrins, bearing a hydroxyl and halide functional group, are privileged building blocks in organic synthesis and could be conveniently converted into other significant organic intermediates such as azido alcohols, amino alcohols, and epoxides, all of which are widely used in the synthesis of many highly value-added chemicals. Up to date, halohydrins could be prepared by halohydrination of olefins, reduction of haloketones, ring-opening reaction of epoxides, and nucleophilic substitution of benzyl halides (Scheme 1, A). In most cases, the halo atom of halohydrins was introduced by the halo cation such as N-halosuccinimides and their analogues which are not good choices for large-scale halogenations because of expensive price and low atomic economy. Inspired by the enzyme-catalyzed aerobic oxidative halogenation in nature, a sophisticated approach by in situ generating the halogenating reagent from oxidant and halide salts is widely applied in halogenations especially in the oxidative halohydroxylation of olefins which provided a direct approach to halohydrins, although solvent, halide source, acid, oxygen source, and oxidant were required in these approaches.

Scheme 1 The synthesis of halohydrins
The β-halogenation of alcohols provides another direct approach to halohydrins. However, the reported β-halogenation of alcohols with halo cations always delivered mixtures of products (Scheme 1, B). As our continuous development of DMSO-based reactions, we herein reported our success in β-halogenation of alcohols with in situ generation of halo cation from sodium halides and DMSO (Scheme 1, C). The magic multiple role of DMSO successfully enabled this novel transformation. Very importantly and interestingly, the diastereoselectivities of this transformation were very high (>25:1)

This reaction began with an unexpected bromination. As reported, the combination of aqueous HBr and DMSO showed high efficiency in the aromatic bromination of 2-naphthol to deliver aryl bromide in 79% yield (Scheme 2, eq 1). To our surprise, when changing the substituent of naphthalene from –OH (1) to –CH(OH)CH₃ (3a), the aromatic bromination was totally suppressed, and aliphatic bromination at the methyl group occurred to afford bromohydrin 4a in 33% yield (Scheme 2, eq 2). Due to the importance of the halohydrins, this bromination drew our great interest.

We then investigated the substrate scope of this novel bromination (Scheme 3). Various benzyl alcohols worked well under the standard conditions. It was noteworthy that when an electron-donating group such as methoxy, tert-butyl, or methyl was contained at the aryl ring, the corresponding bromohydrins 4d, e and 4g were highly selectively obtained in good yields, respectively. The reported bromination on the electron-rich arenes was not detected in this protocol. In addition, heteroarenes such as benzothiophenyl- and benzofuranyl-substituted alcohols were well tolerated in this transformation and converted into bromohydrins 4m, n in moderate yields. Furthermore, the gram-scale reaction of 3a with 69% yield shows the potential application of this low-cost protocol. However, no bromohydrin was detected when alcohols without benzylic substituent was exposed under the standard conditions.

It is very challenging to control the diastereoselectivity of β-functionalization of alcohols. To our delight, the six-membered cyclic alcohols 3o, p produced trans-bromohydrins 4o, p as the sole product in high efficiency (Scheme 4). Although the five-membered and seven-membered cyclic alcohols 3q, r afforded the target bromohydrins 4q, r in moderate yields, the diastereoselectivities of these brominations were also very high (>25:1). The substituents on the phenyl ring had little influence on the efficiency and diastereoselectivity.

The optimized bromination conditions (Table 1). The reaction did not work when changing aqueous HBr to KBr (Table 1, entry 1). This experiment revealed the acidic conditions were essential for this reaction. When HBr was generated by KBr and H₂SO₄ in situ, 4a could be obtained in 38% yield (Table 1, entry 2). The yield increased to 44% using NaBr instead of KBr (Table 1, entry 3). Other organic acids such as TsOH, MsOH, or TfOH showed lower efficiency than NaBr. When NaBr was replaced by KBr (Table 1, entry 4), the yield increased to 44% using H₂SO₄. The amount of acidic additive such as TsOH, MsOH, or TfOH showed lower efficiency than NaBr instead of KBr (Table 1, entry 5). Other organic acids such as KBr and H₂SO₄ showed lower efficiency than NaBr instead of KBr (Table 1, entry 6). This experiment revealed the acidic conditions were essential for this reaction. When HBr was replaced by KBr (Table 1, entry 7). This experiment revealed the acidic conditions were essential for this reaction. When HBr was replaced by NaBr (2 equiv) in DMSO, the yield increased to 46% (Table 1, entry 8). When the reaction was performed with 1.2 equiv of NaBr, only 46% yield of 4a was obtained (Table 1, entry 9). Compound 4a was not detected when the reaction was carried out in other solvents (Table 1, entries 12 and 13). These results indicated that DMSO was indispensable in this bromination.

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Table 1: Optimization of the Reaction Conditions

<table>
<thead>
<tr>
<th>Entry</th>
<th>[Br] (equiv)</th>
<th>Additive (equiv)</th>
<th>Solvent</th>
<th>T (°C)</th>
<th>Yield (%)b</th>
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<tr>
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<td>44</td>
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<tr>
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<td>H₂SO₄ (4)</td>
<td>THF</td>
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</table>

* Reaction conditions: The solution of 3a (0.5 mmol), bromide source, and additive in solvent (1 mL) was stirred under air for 24 h.

B Isolated yields.

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The bromohydrin 4a was conveniently converted into other valuable products (Scheme 6). Azido alcohol 7, the key intermediate of β-blocker pronethalol\textsuperscript{19a} and other bioactive molecules,\textsuperscript{19b} was synthesized in 98% yield by stirring 4a with NaNH\textsubscript{4} at 60 °C in DMSO. Exposure of 4a with aqueous NaOH solvent in THF afforded epoxide 8 in 95% yield, which could easily react with amines to produce amino alcohol drugs.\textsuperscript{19c} Amino alcohol 9 also could be synthesized by treating 4a with ammonium hydroxide in MeOH. The reaction of 4a and CO\textsubscript{2} in the presence of NMe\textsubscript{4}HCO\textsubscript{3} provided carbonic ester 10 in 97% yield.

To investigate the mechanism of this bromination, control experiments were performed. Ethyl naphthalene 11 could not be brominated under the standard conditions, indicating the hydroxyl was indispensable for present bromination (Scheme 7, eq 5). Treatment of alcohol 3a under standard conditions in the absence of NaBr afforded alkene 12 in 39% yield (Scheme 7, eq 6). Subjecting the obtained alkene 12 back to the standard conditions led to 4a in 76% yield, which indicated that alkene 12 might be the key intermediate of this transformation (Scheme 7, eq 7). Previous studies reported that the hydrobromic acids could be oxidized by DMSO to molecular bromine (Scheme 7, eq 8).\textsuperscript{11} If Br\textsubscript{2} was generated, it would readily react with...
the in situ generated alkene 12 to afford a trans-dibrominated product 13. However, exposure of trans-13 under the conditions for 24 h provided the product 4p only in 27% yield and showed much lower diastereoselectivity (trans/cis = 9:1, Scheme 7, eq 9). This experiment demonstrated that the corresponding dibromination was not involved in this process. We therefore suspected that the HBr was oxidized to bromo cation (Br+) which was stabilized by DMSO through coordination (Scheme 7, eq 10).

On the basis of the above experimental results and previous reports,18,20 we proposed the mechanism of this halogenation of alcohols in Scheme 8. Under acidic conditions, alkene 12u is generated by the hydroxyl elimination process. Meanwhile, Br+ is generated and immediately coordinated by DMSO to give (DMSO)nBr+DMS (n = 1–3).10 The electrophilic addition of Br+ to alkene 12u preferentially affords bromonium A rather than B because of the steric hindrance of the methyl group. Further nucleophilic attack of DMSO to A delivers the alkoxy sulfonium C which quickly decomposes to produce trans-bromohydrin 4u.10,16a

In conclusion, we have developed a novel β-halogenation of benzylic alcohols for the efficient synthesis of high-value halohydrins. The simple reaction conditions, easily available reagents, and high diastereoselectivity control make this protocol very attractive and practical. Mechanistic studies reveal that halo cation (X+) rather than molecular halogen is involved in the transformation. This reaction demonstrates a new application of DMSO and HX in organic synthesis and would promote the application of the alkene in situ generation strategy.

Scheme 7

**Scheme 7**

![Scheme 7](image)

**Scheme 8**

![Scheme 8](image)

**Scheme 8** Proposed mechanism

**Funding Information**

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**Supporting Information**

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**References**
