Single-Electron-Transfer Oxidation of Trifluoroborates and Silicates with Organic Reagents: A Comparative Study

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Abstract  In this report, the single-electron-transfer oxidation of alkyl trifluoroborates and silicates has been studied. Different types of oxidation reagents have been examined, focusing on organic oxidants and particularly the use of dyes in photocatalytic oxidations. Both trifluoroborates and silicates could provide C-centered radicals when using a tritylium salt or the Ledwith–Weitz aminium salt. Photocatalysis with the Fukuzumi reagent suggested that trifluoroborates are more easily oxidized than biscatecholato silicates under these conditions.

Key words radicals, dyes, photocatalysis, trifluoroborates, silicates, single-electron transfer, oxidation

The single-electron-transfer (SET) oxidation of soft carbanions is a very versatile method to access to C-centered radicals. Among possible candidates, ate complexes based for instance on boron, trifluoroborates being the most popular reagents, have already shown versatile reactivities for the generation of radicals. To a lesser extent, hypervalent biscatecholato silicon species have recently emerged as very promising alternatives to the boron derivatives, avoiding any release of noxious fluorinated byproducts. Their synthesis is known yielding bench-stable compounds, and their high electron density make them suitable candidates for oxidation. In this letter, we provide new elements on the SET oxidation of alkyl trifluoroborates and silicates, notably focusing on the use of organic oxidants (Scheme 1).

Our own endeavors in this domain started with the copper(II)-mediated oxidation of alkyl trifluoroborates 1 by Kumada and coworkers, a series of alkyl (from primary to tertiary ones) trifluoroborates were engaged in oxidative processes. Postfunctionalization of the resulting radical intermediate was achieved by TEMPO spin trapping, allylation, and conjugate addition.

Following these preliminary reports, we wanted to investigate the use of nonmetallic oxidants. We initially showed that the Dess–Martin periodinane (DMP) could be efficiently used for the oxidation of trifluoroborates. Tritylium tetrafluoroborate, an underexplored oxidant, was also tested with a series of trifluoroborates (Scheme 2). For reasons which need to be elucidated, DMF did not appear as the best solvent for these oxidations. Gratifyingly, good

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Scheme 1 Generation of alkyl radicals by SET oxidation of alkyl trifluoroborates and silicates
Cluster Synlett yields of TEMPO adducts 4 were obtained in Et₂O as solvent with benzyl precursor (4a obtained), but also in secondary (4e and 4g) and primary series (4d,d’). Only tert-butyl precursor 1f failed to give a good yield of product (4f, 25%), presumably for steric reasons. Interestingly, these conditions proved to be compatible with conjugate addition since methyl vinyl ketone (MVK) adduct 5 was isolated in satisfactory yield (63%).

Next, we investigated the reactivity of biscatecholato pentavalent silicates 3. These substrates are amenable to large-scale synthesis and can be rendered rock stable by complexing the potassium counterion by the 18-c-6 crown ether. Benzyl silicate 3a served as a preliminary probe (Scheme 3). It was submitted in Et₂O and DMF to one equivalent of tritylium and aminium. In both solvents, tritylium gave poor yields of 4a (< 20%). However, the use of the aminium salt was more rewarding (86% of 4a in DMF, 16% in Et₂O). This oxidant proved to be competent in DMF for secondary and primary alkyl substrates giving, respectively, 44% of 4h and 61% of 4d,d’. Tritylium can also be used as a reliable alternative oxidant for the silicates 3.

Because of its mild conditions and high substrate tolerance, visible-light photocatalytic oxidation was the obvious next step. In the case of trifluoroborates, several groups have established the feasibility of this transformation using ruthenium(II)- or iridium(III)-based photocatalysts. Of

We also examined the possibility of using Ledwith-Weitz aminium salt (oxidation potential: 1.06 V vs. SCE) as SET oxidative agent of soft carbanions which, to the best of our knowledge, has never been accomplished. A strong solvent effect (Et₂O vs. DMF) was observed in the oxidation of 1a, respectively 2% vs. 69% of 4a. This led us to pursue our study in DMF with this oxidant. However, even in this solvent, the results proved to be much less satisfying compared to the ones obtained with the tritylium oxidant. Only 27% yield (4e) with the secondary substrate 1e, and no TEMPO adduct in the primary alkyl series.

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note, the resulting radicals can be engaged in photoredox/nickel dual catalysis.\textsuperscript{13g–k} While pentafluorosilicates 2 failed in our hands to undergo any oxidation,\textsuperscript{14} we recently showed that biscatecholato silicates constitute advantageous alternatives to the trifluoroborates since they allow upon iridium(III) \{Ir[d(F(CF\textsubscript{3})\textsubscript{3}ppy)\textsubscript{2}(bpy)][PF\textsubscript{6}])\} photocatalysis the generation of very unstabilized primary radicals, also successfully engaged in photoredox/nickel dual catalysis.\textsuperscript{3}

Herein, we wanted to examine the possibility to use organic dyes\textsuperscript{12,15} as possible catalysts for the oxidation of these soft carbanions. Based on their frequent use, the following dyes were considered: eosin Y, fluorescein,\textsuperscript{16} and Fukuzumi acridinium as catalysts.\textsuperscript{17} A preliminary screening with benzyltrifluoroborate 1a showed that the Fukuzumi catalyst was by far the best one (Scheme 4).

![Scheme 4 Photocatalytic oxidation of trifluoroborates and biscatecholato silicates by organic dyes.](image-url)

Similar behavior was observed for 3a. Therefore we kept this catalyst for further testing. Both substrate families showed the same trend, that is, the less stabilized is the generated radical, the lower is the yield. Thus, for trifluoroborates, a gradual decrease of yield was observed from benzyl product 4a to least stabilized primary radical adducts 4d, 4d'. One could argue that 1g, a secondary substrate, should have given a better yield. But in that case, the final radical is a tertiary one which may undergo competitive pathways and lead to only 18% of 4g. In the case of silicates 3, only stabilized benzyl and allyl radicals could be generated (66% for 4a, 31% for 4b). Interestingly, allyl trifluoroborate 1b and allylsilicate 3b provided 4b in close yields (38% vs. 31%). In sharp contrast, however, secondary trifluoroborates could give TEMPO adducts 4e and 4i contrary to secondary silicate 3h (no 4h formed).\textsuperscript{18}
A direct correlation of these findings with redox potentials is not obvious. Oxidation potentials for trifluoroborates span from 1.1 V (benzyl et alkoxy methy) to 1.83 V vs. SCE (primary and aryl)\textsuperscript{1,2} while they have been determined to range from 0.61 V for benzylsilicate to 0.75 V vs. SCE for 3\textsuperscript{d}. Some other key factors are at play in these reactions that we will try to uncover. In all the successful oxidations, TEMPO would act as a sacrificial oxidant to regenerate the photocatalyst and sustain the photocatalytic cycle in agreement with the literature data.\textsuperscript{3,8}

In conclusion, this study shows the unprecedented oxidation of trifluoroborates and silicates with a triptylum or an aminium salt as stoichiometric oxidant to generate C-centered radicals. Photocatalytic oxidation could also be achieved with the Fukuzumi acridinium showing a higher yield of 

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Supporting Information

Supporting information for this article is available online at http://dx.doi.org/10.1055/s-0035-1561337.

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(8) Tritylum is known as a hydride acceptor, for a recent application, see: (a) Xie, Z.; Liu, L.; Chen, W.; Zheng, H.; Xu, Q. H.; Yuan, H.; Lou, H. Angew. Chem. Int. Ed. 2014, 53, 3904; and with 1 equiv of DMP, 26% of 

(9) To a Schlenk flask was added potassium 5-hexenyl-1-trifluoroborate (1d) or potassium [18-crown-6] bis(catecholate)-5-hexenyl-1-silicate (3d. 0.3 mmol, 1 equiv), and TEMPO (0.9 mmol, 141 mg, 3 equiv). The Schlenk flask was sealed with a rubber septum, and evacuated–purged with vacuum–argon three times. Degassed Et2O or DMF (3 mL) was introduced followed by two freeze–pump–thaw cycles. The reaction mixture was stirred at room temperature for 24 h under an argon atmosphere. The reaction mixture was diluted with Et2O (50 mL), washed with H2O or NaHCO3 (2×), and with 1 equiv of DMP, 26% of 


(11) In comparison, oxidation with 1 equiv of Cu(OAc)2 gave 45% yield of 


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