Thiophosgene

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Introduction

Thiophosgene, also known as thiocarbonyl chloride, is a red-orange liquid with a strong and suffocating odor. Thiophosgene is irritant to eyes (lachrymatory), skin and respiratory tracts; nevertheless, it is less toxic than phosgene. Thiophosgene reacts with nucleophilic sites in a number of functional groups like amines, alcohols, phenols, thiols, oximes, among others. A variety of heterocycles can be constructed when there are two nucleophiles close. As a consequence, thiophosgene has many applications in organic synthesis. It can be used to prepare isothiocyanates which serve as important scaffolds to provide compounds such as thioureas, thiazoles\(^1\)\(^2\) or thiocarbamates.\(^3\) Besides, thiones,\(^4\) oxadiazolones,\(^5\) thiocarbonates,\(^6\) chlorothioformates,\(^7\) and heptathiodicarbonates\(^8\) can be synthetized with the employment of this reagent.

Thiophosgene was first prepared in small amounts in 1870 by Rathke;\(^9\) then in 1887, Klason\(^10\) developed a more efficient methodology via reduction of trichloromethane-sulfenyl chloride with zinc.

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<th>Use of Thiophosgene</th>
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<td>(A) Recently, in the search of novel selective antitumor agents with more selectivity to destroy malignant cells, a group of optically active amines 1 were converted into isothiocyanates 2 using a solution of thiophosgene. Isothiocyanates were used as intermediates in the synthesis of a series of active thioureas and 2-aminobenzothiazoles evaluated in biological tests.(^1)</td>
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<td>(B) Figadère and co-workers(^4) described the formation of a group of achiral 5,5-disubstituted N-acetylazoxazolidine-2-thiones 4 using different amino alcohols 3, thiophosgene and Et(_3)N in THF and subsequent acetylation on the nitrogen. The yields reported in the first step range from 20 to 44%. These heterocycles were studied as achiral auxiliaries in the C-glycosylation of lactol acetates.</td>
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<td>(C) In 2011, Rozen's group(^6) reported a novel methodology for preparing aromatic difluoromethylene dioxides 8 from deactivated or mildly deactivated aromatic rings (5,6) using bromine trifluoride (BrF(_3)). In order to introduce the fluorides without radical side reactions, they rationalized the inclusion of sulfur as a soft base using cyclic thiocarbonates 7 as intermediates, which were obtained by the reaction of aromatic derivatives with thiophosgene in 90–95% yields.</td>
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<td>(D) Hagooly et al.(^7) described a methodology to afford ROCF(_2)Cl ethers 10, avoiding the formation of trifluromethyl ethers as byproducts. First, the alcohol dissolved in Et(_3)N reacted with a solution of thiophosgene forming the respective chlorothioformates 9 in 80–95% yields. Then, the reaction of BrF(_3) with chlorothioformates provided the desired products in short reaction times.</td>
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(E) In 2013, Weber et al. described the utility of thiophosgene in the synthesis of 1,3-dithietane rings in the search for a new pathway to form heptadiithiocarbonates as potential RAFT reagents. The potassium salts 11, 14 and 17 were treated with thiophosgene resulting in the formation of products characterized as 13, 16 and 19. Either the yields of the global reaction or the isolation of intermediates (12 and 18) depended on the solvent and the mode of addition of thiophosgene.

(F) In the synthesis of BODIPYs, Wang et al. used the formation of dipyrrylketones 22 as the key step; such compounds were obtained in two stages: first, the reaction of pyrrole derivatives 20 and thiophosgene to get the corresponding dipyrrylthioketones 21 in 40 and 43% yield and then a subsequent oxidative hydrolysis.

(G) With the aim of finding new active molecules as potent HIV-1 TR inhibitors, Monforte and co-workers designed a synthesis of a series of novel benzimidazolones and analogues. In this group of compounds, a derivative with a thio carbonyl moiety 24 was synthesized; this product was obtained by the reaction of 23 with thiophosgene in acetone in 14% yield and showed high inhibitory potency.

References
