Study of Ring-Opening Reaction of Spiro[5.2]octenes with Aqueous Hydrohalic Acid: Substituent Effect on the Regioselectivity

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Abstract: We describe here the regioselective ring-opening reaction of spiro[5.2]octenes with hydrohalic acids. It was observed that the electronic nature of a substituent on the cyclopropane ring would control the regioselectivity.

Key words: alkenycyclopropanes, electrophilic addition, regioselectivity, ring opening, spiro compound

Cyclopropane framework is an attractive organic motif in terms of its biological property as well as its unique chemical reactivity. The relief of the high ring-strain energy, which provides a potent enthalpic driving force, facilitates a variety of unique chemical transformations.1 Cyclopropane compounds are found as naturally occurring substances such as illudins,2b,3 duocarmycins,2b,4 and their related compounds, which display potent antitumor activity (Scheme 1). The bioorganic studies of these natural products proved that the ring-opening reaction of the cyclopropane ring induced by a nucleophilic attack of nucleobases or proteins is a key event.5,6 The mechanism of biological action of illudins is illustrated in Scheme 1.5a Michael addition by a thiol nucleophile of a protein, such as glutathione reductase (GSH), under acidic conditions (pH ~6) gives active intermediate A. The electrophilic intermediate A smoothly reacts with an appropriate nucleophile. As the result, aromatic bioconjugate B is produced by concurrent ring opening of the cyclopropane moiety by the S$_2$2'-type reaction.

Based on the mechanism, extensive efforts have been devoted to designing the illudin and duocarmycin analogues and to understanding their biological properties.6–8 Barone and co-workers examined the ring-opening reaction of duocarmycin SA derivatives, whose spirocyclopropane ring is electrophilically activated by the conjugation with a carbonyl group.9 They made clear that a nucleophile attacks preferably at the least substituted carbon atom of the cyclopropane ring. Moreover, the computational studies revealed that both the electronic and the steric effects of the substituent can influence the reactivity of the ring-opening reaction.9 Therefore, they concluded that the substituent effect would be helpful in exploring the chemical transformations of duocarmycin SA as well as in designing new analogues.

Recently, we reported the synthesis of spirocyclic alkenycyclopropanes 1, bearing a substituent on the cyclopropane ring, from fused cyclobutanols.10 The spirocyclic core of 1 is structurally related to illudins. Our intention is to design a new antitumor DNA-alkylating agent by using our synthetic reaction. For this purpose, the ring-opening reaction of alkenycyclopropane 1, whose cyclopropane is not conjugated with an electron-withdrawing group, must be examined under aqueous conditions. Additionally, investigation of the regioselectivity would be an important issue because the cationic intermediate C has three different reactive sites with a nucleophile X$^\text{−}$ (Scheme 2). When nucleophilic (X$^\text{−}$) attack occurs at the substituted or non-substituted cyclopropane carbon atoms, 1,5-adducts 2 or 3, respectively, will be obtained via a conjugate addition (modes a and b). On the other hand, direct addition of X$^\text{−}$ at the tertiary cationic carbon of C will give 1,2-adduct 4 (mode c). Although a number of transition-metal-catalyzed ring-opening reactions of alkenycyclopropanes were reported,11e,11f to the best of our knowledge, only limited investigations for acid-mediated ring-opening reactions of inactivated alkenycyclopropanes can be found in the literature.10,11 Moreover, the regioselective issues are unexplored. Herein, we describe the regioselectivity in the ring-cleavage reaction of spiro[5.2]octenes 1 with various aqueous hydrohalic acids (HX).
At the outset of our study, ring-opening reaction of 1a\(^9\) bearing an alkyl substituent under aqueous acidic conditions was examined (Table 1). When 1a was treated with aqueous hydroiodic acid (HI), which was prepared in situ from TMSCl, NaI, and H\(_2\)O (2 equiv each)\(^1\) in acetonitrile at ambient temperature, ring-opening reaction proceeded smoothly giving 2aa in 68% (Table 1, entry 1). Namely, the product 2aa was formed through the nucleophilic substitution at the more substituted carbon of 1a (mode a in Scheme 2). No isomeric adduct such as 3aa, which will be produced by the C–X bond formation at the less substituted carbon (mode b in Scheme 2), was observed. As is the case in HI, secondary alkylbromide 2ab was obtained as a single isomeric product in the reaction of 1a with aqueous hydrobromic acid (HBr, Table 1, entry 2). The same regioselectivity as in acetonitrile was also observed in dichloromethane, although the chemical yield was decreased (Table 1, entry 3). With hydrochloric acid (HCl), only electrophilic addition of Cl\(^–\) to the substituted carbon occurred to afford 2ac in 86% yield (Table 1, entry 4). It was found that, in the ring-opening reaction of 1a with HX, a nucleophile attacks regioselectively at the more substituted carbon regardless of the nature of the nucleophiles.

The regioselectivity of the ring-opening reaction can be explained as shown in Scheme 3. After protonation of the alkene moiety of 1, the corresponding tertiary cation intermediate C is generated. Major product 2 should be obtained via nucleophilic attack of X\(^–\) at the more substituted carbon [transition-state structure (TS) D], and minor product 3 via nucleophilic attack at the less substituted carbon (TS E). Partial positive charge (δ\(^+\)) is delocalized at the tertiary or secondary carbon in D or E, respectively.\(^1\)\(^4\) In the reactions of alkyl-substituted substrate 1a, TS D would be more stable than TS E owing to the cationic stability (primary vs. secondary carbocation), resulting in product 2a with high regioselectivity.

Next the reactions of 1b bearing an ether substituent with HX were carried out (Table 2). When 1b was treated with HI, formation of primary alkyl iodide 3ba was observed as a minor product along with secondary alkyl iodide 2ba (Table 2, entry 1). In the reaction of 1b with HBr and HCl, the mixture of 2b and 3b were also obtained, respectively (Table 2, entries 2 and 3). We guessed that an inductive effect would rationalize the formation of 3b.\(^1\)\(^5\) Namely, the partially delocalized positive charge at the substituted carbon in TS D from 2b would be somewhat destabilized by the ether moiety (negative inductive effect). As the result, the regioisomeric ring-opening reaction via TS E giving 3b also proceeds as a minor path.

In order to make clear the negative inductive effect by the side chain on the cyclopropane ring, we investigated the reaction of 1c having an electron-withdrawing ester substituent (Table 3). Treatment of 1c with HI in acetonitrile afforded no 2ca but primary alkyl iodide 3ca in 75% yield (Table 3, entry 1). The regioselectivity decreased when the reaction of 1c was carried out in the less polar solvents (Table 3, entries 2 and 3). The solvent effect supports that the regioselectivity of the ring-opening reaction is dependent on the electronic factor of the intermediate or transition state. When the reaction of 1c was carried out with HBr, primary alkyl bromide 3cb was obtained along with

![Scheme 2](image)

**Scheme 2** Possible regioisomeric adducts 2–4 in the reaction of 1 with HX

<table>
<thead>
<tr>
<th>Entry</th>
<th>HX</th>
<th>Solvent</th>
<th>Product ratio</th>
<th>Yield (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>HI</td>
<td>MeCN</td>
<td>2aa/3aa = 100:0</td>
<td>68</td>
</tr>
<tr>
<td>2</td>
<td>HBr</td>
<td>MeCN</td>
<td>2ab/3ab = 100:0</td>
<td>67</td>
</tr>
<tr>
<td>3</td>
<td>HBr</td>
<td>CH(_2)Cl(_2)</td>
<td>2ab/3ab = 100:0</td>
<td>49</td>
</tr>
<tr>
<td>4</td>
<td>HCl</td>
<td>MeCN</td>
<td>2ac/3ac = 100:0</td>
<td>86</td>
</tr>
</tbody>
</table>

\(^a\) Conditions: 1a (0.1 mmol), HX (1.5 or 2.0 equiv), MeCN (0.2 M), r.t., 1–3 h.

\(^b\) Compound 1a was recovered in 12% yield.

![Scheme 3](image)

**Scheme 3** Plausible mechanism of regioselectivity

HI, formation of primary alkyl iodide 3ba was observed as a minor product along with secondary alkyl iodide 2ba (Table 2, entry 1). In the reaction of 1b with HBr and HCl, the mixture of 2b and 3b were also obtained, respectively (Table 2, entries 2 and 3). We guessed that an inductive effect would rationalize the formation of 3b.\(^1\)\(^5\) Namely, the partially delocalized positive charge at the substituted carbon in TS D from 2b would be somewhat destabilized by the ether moiety (negative inductive effect). As the result, the regioisomeric ring-opening reaction via TS E giving 3b also proceeds as a minor path.
The regioisomeric adduct 2cb (Table 3, entry 4). The observation indicates that the regioselectivity would be also dependent on the nature of the nucleophile (X–). The reaction with HCl gave considerably complicated results (Table 3, entries 5 and 6). At ambient temperature, ring-opening products 2cc and 3cc were obtained in 25% yield with a 33:67 diastereomeric ratio. As a major product, cyclopropylcarbinol 4cd was obtained in 32% yield as a 1:1 diastereomeric mixture. Formation of 4cd resulted in 1,2-addition (mode c in Scheme 1) of water. As minor products, alcohols 2cd (R = CH2OBz, X = OH) and 2cd′ (R = OH, X = CH2OBz) were also isolated in 8% and 2% yield, respectively. When the reaction was carried out at 80 °C, no 1,2-adduct 4 was detected but ring-opening products 2cc, 3cc, 2cd, and 2cd′ were produced. It indicates that 4cd is a kinetic product, and 4cd is transformed into conjugated adducts at higher temperature.

The regioselective formation of 3ca from ester 1c strongly supports that the negative inductive effect of the substituent would control the position of the ring cleavage. Namely, the delocalized cation in the TS D (Scheme 3) would be delocalized by the negative inductive effect of the alkoxycarbonyl substituent. As the result, the ring-opening reaction preferably proceeds via TS E. Production of alcohol 2cd′ (Table 3, entries 5 and 6) can be explained by the formation of oxonium intermediate F,16 which would be generated by an intramolecular cyclopropane ring-opening reaction as shown in Scheme 4. As chloride anion (Cl–) is less reactive than bromide and iodide anions,17 the nucleophilic attack to intermediate C by the intramolecular carbonyl group or external water molecule occurred faster than chlorination.

The ring-opening reaction of 6-oxo analogue 5 under acidic conditions was also examined (Scheme 5). Treatment of 5 with TMSCl and NaI in aqueous acetonitrile smoothly afforded enone 6 in 83% yield even at 0 °C. Product 6 was obtained as a 1:1 diastereomeric mixture but no formation of its regioisomeric product was observed. The nucleophile (I–) exclusively attacks at the less

Table 2 Reaction of 1b Having a Benzylxymethyl Substituent with Various HXa

<table>
<thead>
<tr>
<th>Entry</th>
<th>HX</th>
<th>Product ratio</th>
<th>Yield (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>HI</td>
<td>2ba/3ba = 71:29</td>
<td>65</td>
</tr>
<tr>
<td>2</td>
<td>HBr</td>
<td>2bb/3bb = 59:41</td>
<td>88</td>
</tr>
<tr>
<td>3</td>
<td>HCl</td>
<td>2bc/3bc = 72:28</td>
<td>78</td>
</tr>
</tbody>
</table>

a Conditions: 1b (0.1 mmol), HX (1.5 or 2.0 equiv), MeCN (0.2 M), r.t., 0.5–3 h.

Table 3 Reaction of 1c Having a Benzoyloxymethyl Substituent with Various HXa

<table>
<thead>
<tr>
<th>Entry</th>
<th>HX</th>
<th>Solvent</th>
<th>Product ratio</th>
<th>Yield (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>HI</td>
<td>MeCN</td>
<td>2ca/3ca/4ca = 0:100:0</td>
<td>75</td>
</tr>
<tr>
<td>2</td>
<td>HI</td>
<td>CH3Cl</td>
<td>2ca/3ca/4ca = 9:91:0</td>
<td>64</td>
</tr>
<tr>
<td>3</td>
<td>HI</td>
<td>toluene</td>
<td>2ca/3ca/4ca = 38:62:0</td>
<td>69</td>
</tr>
<tr>
<td>4</td>
<td>HBr</td>
<td>MeCN</td>
<td>2cb/3cb/4cb = 23:77:0</td>
<td>73</td>
</tr>
<tr>
<td>5</td>
<td>HCl</td>
<td>MeCN</td>
<td>2ce/3ce/4ce = 27:24:49b</td>
<td>67</td>
</tr>
<tr>
<td>6c</td>
<td>HCl</td>
<td>MeCN</td>
<td>2ce/3ce/4ce = 72:28:0d</td>
<td>82</td>
</tr>
</tbody>
</table>

a Conditions: 1c (0.1 mmol), HX (1.5 or 2.0 equiv), solvent (0.2 M), r.t., 1–3 h.
b Compounds 2cc, 2cd, 2cd′, 3cc, and 4cd were obtained in 8%, 8%, 2%, 17%, and 32%, respectively.
c The reaction was carried out at 80 °C.
d Compounds 2cd and 2cd′ were also obtained in 29% and 14% yield, respectively.
substituted cyclopropane carbon atom of 5 as is the case of 1c.

In conclusion, we have examined the ring-opening reaction of spiroyclic alkenylcyclopropanes under aqueous acidic conditions. It was elucidated that the regioselectivity in the cyclopropane cleavage was significantly dependent on the electronic nature of the substituent on the cyclopropane ring. The nature of the nucelophile also slightly influences the regioselectivity of the ring-opening reaction. We are currently engaging in the synthesis of new DNA-alkylating spiroyclic cyclopropane compounds. The new findings on the regioselectivity would help us in molecular designing.

Typical Procedure for the Ring-Opening Reaction
To a mixture of 1 (0.10 mmol), NaI (0.20 mmol), and H2O (0.20 mmol) in MeCN (0.2 M) was added TMSCl (0.20 mmol). Otherwise, to a mixture of 1 (0.10 mmol) in MeCN was addedaq concd HCl or HBr (0.20 mmol). After being stirred for 1 h at an appropriate temperature, the reaction mixture was quenched with sat. Na2SO4. The aqueous layer was extracted twice with EtOAc. The combined organic layers were washed with brine, dried over Na2SO4 and concentrated in vacuo. The residue was purified by silica gel chromatography (hexane–EtOAc) to afford desired

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References


